

## HEAT AND THERMODYNAMICS NOTES FOR FSC%0A

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Revision Notes on Thermodynamics Thermodynamics:- It is the branch of physics which deals with process involving heat, work and internal energy. Thermodynamics is concerned with macroscopic behavior rather than microscopic behavior of the system.

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Chapter: 11 (Heat and Thermodynamics) In this chapter, we will study the following important questions: Describe the fundamental postulates (or assumptions) of the kinetic theory of gases.

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Physics Notes Class 11 CHAPTER 12 THERMODYNAMICS The branch dealing with measurement of temperature is called thermometry and the devices used to measure temperature are called thermometers. Heat is a form of energy called thermal energy which flows from a higher temperature body to a lower temperature body when they are placed in contact. Heat or thermal energy of a body is the sum of

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### **Lectures on Heat and Thermodynamics Galileo**

5 The altar is a large airtight box, with a pipe leading from it to another enclosed container filled with water. When the fire is set on top of the altar, the air in the box heats up and expands into a second  
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The branch of science which deals with the quantitative relationship between heat and other forms of energies is called thermodynamics. Some Important Terms Related to Thermodynamics (i) System It refers to the part of universe in which observations are carried out.

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### **Notes on Thermodynamics The key idea is that**

Notes on Thermodynamics The topic for the last part of our physics class this quarter will be thermodynam-ics. Thermodynamics deals with energy transfer processes. The key idea is that materials have "internal energy". The internal energy is the energy that the atoms and molecules of the material possess. For example, in a gas and liquid the molecules are moving and have kinetic energy. The

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### **Thermodynamics Notes UC Santa Barbara**

Heat capacity C describes the amount of heat required to change the temperature of a substance:  $C =$  By definition, the heat capacity of water at 15 C is 1 cal K<sup>-1</sup> g<sup>-1</sup> or 18 cal K<sup>-1</sup> mol<sup>-1</sup> (i.e., the heat required to heat 1 gram of water from 14.5 to 15.5 C is 1 calorie).

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